

Ruthenium Complexes with Vinyl, Styryl, and Vinylpyrenyl Ligands: A Case of Non-innocence in Organometallic Chemistry

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Abstract: We herein describe a systematic account of mononuclear ruthenium vinyl complexes $L-\{Ru\}-CH=CH-R$ where the phosphine ligands at the $(PR'_3)_2Ru(CO)Cl=\{Ru\}$ moiety, the coordination number at the metal ($L = 4$ -ethylisonicotinate or a vacant coordination site) and the substituent R ($R = n$ -butyl, phenyl, 1-pyrenyl) have been varied. Structures of the enynyl complex $Ru(CO)Cl(PPh_3)_2(\eta^1:\eta^2-nBuHC=CHC\equiv C^oBu)$, which results from the coupling of the hexenyl ligand of complex **1a** with another molecule of 1-hexyne, of the hexenyl complexes $(nBuCH=CH)Ru(CO)Cl(P^iPr_3)_2$ (**1c**) and $(nBuCH=CH)Ru(CO)Cl(PPh_3)_2(NC_5H_4COOEt-4)$ (**1b**), and of the pyrenyl complexes $(1-Pyr-CH=CH)Ru(CO)Cl(P^iPr_3)_2$ (**3c**) and $(1-Pyr-CH=CH)Ru(CO)Cl(PPh_3)_3$ (**3a-P**) have been established by X-ray crystallography. All vinyl complexes undergo a one-electron oxidation at fairly low potentials and a second oxidation at more positive potentials. Anodic half-wave or peak potentials show a progressive shift to lower values as π -conjugation within the vinyl ligand increases. Carbonyl band shifts of the metal-bonded CO ligand upon monooxidation are significantly smaller than is expected of a metal-centered oxidation process and are further diminished as the vinyl $CH=CH$ entity is incorporated into a more extended π -system. ESR spectra of the electrogenerated radical cations display negligible g -value anisotropies and small deviations of the average g -value from that of the free electron. The vinyl ligands thus strongly contribute to or even dominate the anodic oxidation processes. This renders them a class of truly "non-innocent" ligands in organometallic ruthenium chemistry. Experimental findings are fully supported by quantum chemical calculations: The contribution of the vinyl ligand to the HOMO increases from 46% (Ru-vinyl delocalized) to 84% (vinyl dominated) as R changes from n -butyl to 1-pyrenyl.

Introduction

Redox active ligands that actively participate in electron-transfer processes, in conjunction with appropriate redox active transition metal entities, render an assignment of redox states ambiguous.¹ Such behavior has been denoted as "non-innocent".² When the energy barrier for intramolecular electron transfer is low, charge and unpaired spin are delocalized over the metal ion and the "non-innocent" ligand(s). This may either result in metal stabilized radicals^{3–5} or in the stabilization of

unusual (formal) oxidation states at the metal as, for example, in compound I of cytochrome P-450.⁶ If, on the other hand, the energy barrier of intramolecular electron transfer is sufficiently high, "non-innocent" behavior gives rise to the phenomenon of valence tautomerism.^{7–10} The charge distribution within valence tautomers can be controlled by exterior physical stimuli such as pressure, light, or applied magnetic field.^{11,12} Such behavior sets the ground for the development of molecule based,

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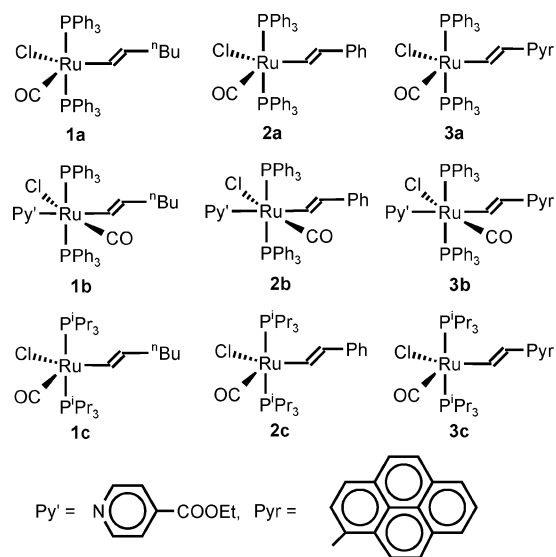
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switchable materials with possible applications, inter alia, for data storage.

Set against this background there is a continued need to prepare and identify redox-active, potentially non-innocent ligands and to study their interaction with various transition-metal moieties. Important sets of redox-active ligands are based on tetracyanoethene (TCNE), 7,7,8,8-tetracyano-*p*-quinodimethane (TCNQ),¹³ or dithiolenes,^{14,15} dioxolenes,^{9,16} diimines,¹⁵ and ligands with mixed donor atoms. While most of the work in this area has concentrated on inorganic coordination compounds, ligands like TCNE and TCNQ may equally well coordinate organometallic entities.^{17–22} Potentially redox-active ligands that form metal–carbon bonds for use in organometallic chemistry are, however, much less well preceded. Examples are paramagnetic alkene complexes of rhodium and iridium,⁴ and manganese complexes containing π -coordinated quinones (or semiquinones or catecholates, respectively).^{23–25} While individual building blocks are organometallic in nature, the quinone complexes still rely on the traditional metal–dioxolene chelate linkage for their interesting electronic properties. The radical derived from a triphenylmethyl substituted vinylferrocene constitutes a rare example of a genuine organometallic valence tautomeric system with a temperature-dependent equilibrium between $\text{Fc}-\text{CH}=\text{CH}-\text{CPh}_3^*$ and $\text{Fc}^+-\text{CH}=\text{CH}-\text{CPh}_3^-$ -states.²⁶

We have recently investigated the anodic electrochemistry of divinylphenylene-bridged diruthenium complexes $\{(\text{PR}_3)_2(\text{CO})\text{Cl}(\text{L}')\text{Ru}\}_2\{\mu\text{-C}_6\text{H}_4(\text{CH}=\text{CH})_{2-1,3}$ or $-1,4\}$ ($\text{R} = \text{Ph}$, $\text{L}' = 4$ -substituted pyridine; $\text{R} = \text{Pr}$, $\text{L}' = \text{none}$)^{27,28} and studied their radical cations and dications by UV–vis–NIR, IR, and ESR spectroscopy. These studies indicated bridge-dominated oxidation processes although the oxidation potentials of ruthenium complexes are intrinsically lower than those of comparable donor substituted phenylene vinylenes. The metal *versus* vinyl ligand contribution to the frontier molecular orbitals (FMOs) depends on the energy gap between the appropriate metal d levels and the π orbitals of the vinyl ligand. Raising the energy of the ruthenium d manifold or decreasing the energy of the ligand π orbitals enhances the metal character of the FMOs.

Chart 1. The Complexes Employed in This Study



Thus, butadienediyl-bridged diruthenium complexes constitute completely delocalized organometallic π -chromophores that defy any assignment of the redox processes as metal or ligand based.²⁹ To further develop this chemistry toward true organometallic valence tautomers and to extended chainlike arrays with high electron mobility along the main chain, it is mandatory to understand and to quantify the principal factors that govern the electron distribution within the ruthenium vinyl entities. Dinuclear systems present the added complexity of possible metal–metal interactions across the bridging ligands. We therefore turned to the mononuclear complexes shown in Chart 1. These were designed such as to allow us to study how (i) extending the π -system of the vinyl ligand, (ii) manipulating the electron density at the metal atom, and (iii) varying the degree of coordinative saturation at the metal atom affect the bonding, the anodic behavior, and metal versus ligand contribution to the redox-orbitals.

Results and Discussion

Syntheses and Characterization of the Vinyl Complexes.

Ruthenium vinyl complexes $(\text{RCH}=\text{CH})\text{Ru}(\text{CO})\text{Cl}(\text{PR}_3)_2$ are easily prepared from hydride complexes $\text{HRu}(\text{CO})\text{Cl}(\text{PR}_3)_n$ ($\text{PR}_3 = \text{PPh}_3$, $n = 3$; $\text{PR}_3 = \text{P}^i\text{Pr}_3$, PCy_3 , $n = 2$) via hydorruthenation of terminal alkynes. Metal hydride insertion into the $\text{C}\equiv\text{C}$ bond occurs in a regio- and stereospecific manner and results in the anti-Markovnikov product where the metal atom is connected to the terminal carbon atom and the remaining substituent is in a trans position at the vinyl group. Depending on the reaction conditions $\text{HRu}(\text{CO})\text{Cl}(\text{PPh}_3)_3$ either gives coordinatively saturated six-coordinated tris(phosphine) derivatives or coordinatively unsaturated five-coordinated square pyramidal bis-(phosphine) complexes with a vacant coordination site trans to the apical vinyl ligand.^{30–32} Irrespective of five- or six-

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coordination, PPh₃ substituted vinyl complexes readily react with sterically less demanding two electron ligands such as CO, isonitriles, amines, or pyridines to give octahedral 18 valence electron complexes with the newly introduced donor opposite to the vinyl ligand. The 16 valence electron complexes (RCH=CH)Ru(CO)Cl(PⁱPr₃)₂, on the other hand, are less prone to addition of a pyridine ligand but may still add CO, isonitriles, or, under simultaneous chloride substitution, a bidentate monoanionic ligand such as acetate.^{33,34}

The reaction of HRu(CO)Cl(PPh₃)₃ with 1-hexyne gave an equilibrium mixture of the five- and six-coordinated bis- and tris(phosphine) hexenyl complexes. Variable temperature NMR spectroscopy in CD₂Cl₂ indicates that the equilibrium lies completely on the side of the red 16 valence electron complex and free phosphine at room temperature, while only six-coordinated (ⁿBuCH=CH)(PPh₃)₃(CO)ClRu is present at or below 203 K or in the solid. Similar findings were just reported for the styryl complex **2a**.³²

When this equilibrium mixture was reacted with isonicotinate, a mixture of the desired complex **1b**, unreacted **1a**, and the disubstitution product [(PPh₃)₂(CO)(NC₅H₄COOEt)₂(ⁿBuCH=CH)Ru]⁺ Cl⁻ was formed from which no pure compound could be isolated. Complex **1b** was finally prepared from HRu(CO)Cl(PPh₃)₂(NC₅H₄COOEt). Like in the case of HRu(PⁱPr₂Ph)₂(CO)Cl(py)³⁵ this latter hydride was obtained as a 78:22 mixture of two mutually interconverting isomers that differ with respect to the arrangement of the four different ligands in the equatorial plane. According to ¹H ROESY the major isomer has the isonicotinate cis to the hydride. For the major isomer slow rotation of the substituted pyridine at room temperature is evident from the four different pyridine resonance signals. Cooling to 193 K is required for the minor isomer where the isonicotinate ligand is trans to the hydride in order to freeze pyridine rotation (*T*_c ≈ 208 K). This large difference in rotational barriers may arise from hydrogen bonding between the ortho proton of the pyridine and the hydride ligand. The inequivalence of the 2,6- and 3,5-hydrogen atoms in both isomers points to static structures where the pyridine planes avoid eclipsing the PRuP axis. Similar observations have been made for pyridine adducts of HRu(CO)X(PⁱBu₂Me)₂ (X = monoanionic ligand) where the pyridine ligand is in a trans position to the hydride.^{36,37} Treating this isomeric mixture with 1-hexyne results in the exclusive formation of **1b**. Under gentle warming HRu(CO)Cl(PPh₃)₃ reacted with excess 1-hexyne to give the dodec-5-en-7-ynyl complex Ru(CO)Cl(PPh₃)₂(η¹:η²-ⁿBuHC=CH-C≡CⁿBu), which was studied by X-ray crystallography (vide infra). Closely related butenynyl complexes have already been reported.³⁸

All complexes of this study (Chart 1) were investigated and characterized by the usual spectroscopic and analytical methods including multinuclear NMR, IR, and electronic spectroscopy. Characteristic features in their ¹H NMR spectra are the two

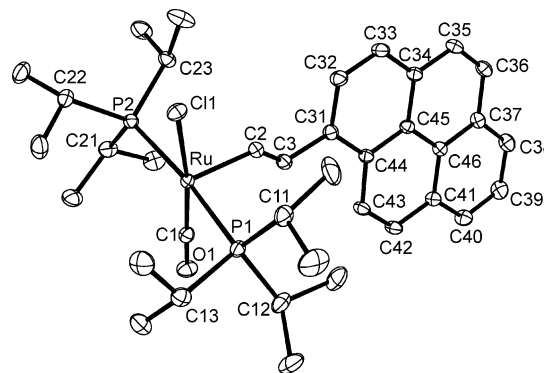


Figure 1. ORTEP representation of complex **3c**·CHCl₃ at a 50% probability level.

doublets for the vinyl protons. The signal of the vinyl proton at the metal-bonded carbon atom C_α resonates at lower field with δ values ranging from 9.26 to 6.90 ppm. It is less sensitive to the ligand environment than the one at the neighboring carbon atom C_β where the shift ranges from 8.79 to 4.76 ppm. The vinyl carbon atoms resonate at 156.4–138.4 ppm (C_α) and 139.6–130.7 ppm (C_β), respectively. Addition of the substituted pyridine ligand to complexes **1a**, **2a**, and **3a** shifts the resonance signals of the vinylic protons and those of the RuCH=CH and the RuCH=CH carbon atom to lower field and that of the ³¹P resonance to higher field. In their IR spectra the vinyl complexes show an intense band of the CO stretch. The position of this band is more sensitive to the number and nature of the phosphine ligands than to that of the vinyl substituent. Thus, in every series **1a–c**, **2a–c**, and **3a–c** the CO band shifts follow the sequence **1a**, **2a**, **3a** > **1b**, **2b**, **3b** > **1c**, **2c**, **3c** such that the five-coordinated PⁱPr₃ substituted complexes feature the most electron-rich metal centers.

X-ray Structures of 1b, 1c, 3c·CHCl₃, 3a-P·2CH₂Cl₂, and of Ru(CO)Cl(PPh₃)₂(η¹:η²-ⁿBuHC=CHC≡CⁿBu). Crystals of the five-coordinated vinyl complex **1c** were grown from methanol, while **3c** crystallizes from chloroform as a monosolvate. Figure 1 provides a plot of the complex molecule of **3c**·CHCl₃ along with the atomic numbering scheme. A similar plot of **1c** is given as Figure S2 of the Supporting Information. In each of the two structures the ruthenium atom adopts the distorted square pyramidal coordination with the strongly trans influencing vinyl ligand in the apical position that is familiar of M(CO)Cl(PR₃)₂X (X = H, vinyl, alkyl, aryl) complexes.^{33,39–42} Distortions arise from a compression of the P–Ru–P angle toward the vacant coordination site to 168.63(4)° (**1c**) or 173.72(1)° (**3c**) and a slight displacement of the metal atom by 0.152(1) Å (**3c**) or 0.198(1) Å (**1c**) out of the best plane defined by the four donor atoms at the pyramid base. The vinyl ligand resides in the plane defined by the Ru, Cl, and C(1) atoms as it is shown by the torsional angles Cl(1)–Ru–C(2)–C(3) of –177.5(5) or 177.7(1)° for **1c** or **3c**, respectively. The angles Ru–C2–C3–C4 of 176.2(5) (**1c**) and Ru–C2–C3–C31 of 178.6(1)° (**3c**) signal hardly any twisting of the vinyl double bond.

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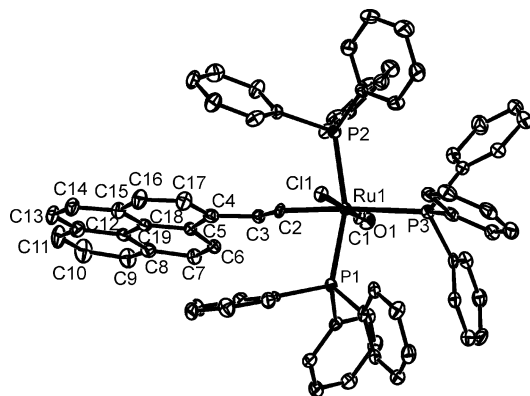


Figure 2. ORTEP representation of one of the two independent molecules of **3a-P** at a 50% probability level.

The pyrenyl part of **3c**, while strictly planar, is rotated out of the Ru–C2–C3 plane by about 25° such that the vinylpyrenyl plane intersects the P–Ru–P axis at an angle of about 67°. The Ru–C2, C2–C3, and C3–C31 bonds of 1.9866(18), 1.350(2), and of 1.473(2) Å are typical of a Ru–C(sp²) single bond, a C=C double bond and a C(sp²)–C(sp²) single bond, respectively. Similar bond lengths for Ru–C2 (1.984(5)) and C2–C3 (1.346(9)) are found for **1c** and related five-coordinated vinyl complexes.^{33,39–42} Structural features of the vinylpyrenyl part of **3c** show essentially the same bond-length pattern as that observed in styryl⁴³ or alkynyl substituted pyrenes.^{44,45} The chloroform solvate molecule forms an almost linear CH⋯Cl–Ru array with a H⋯Cl distance of 2.56 Å, a C⋯Cl distance of 3.551 Å, and a CH⋯Cl angle of 169°.

All attempts to crystallize the PPh₃ substituted counterpart **3a** only led to the formation of few small crystals of a distinctly lighter, yellow color as that of the bulk sample. These were subjected to X-ray analysis and turned out as the PPh₃ adduct **3a-P**. The structure of **3a-P**, as depicted in Figure 2 (only one of two independent molecules is shown), bears clear witness to the distortion of the coordination sphere around ruthenium induced by the strong trans influence of the vinyl ligand and the steric hindrance of three meridionally displaced PPh₃ ligands. Thus, the distances between the ruthenium atom and the mutually trans disposed P atoms P1, P2 (molecule A) or P4, P5 (molecule B) of 2.4221(13) and 2.4024(13) Å or 2.4397(13) and 2.4446(13) Å are distinctly smaller than those of 2.5850(13) or 2.5638(14) Å involving the PPh₃ ligand trans to the pyrenylvinyl group. Steric congestion at the metal center is also evident from the strongly compressed P1–Ru–P2 (P4–Ru–P5) angle of 161.74(4)° (163.83(5)°) and from the bending of the phosphines toward the pyrenylvinyl ligand as opposed to the five-coordinated complexes **1c** and **3c**. This distortion surpasses that observed in similar vinyl complexes bearing PEt₃ or PMe₃ coligands^{46–49} and also that observed in the closely

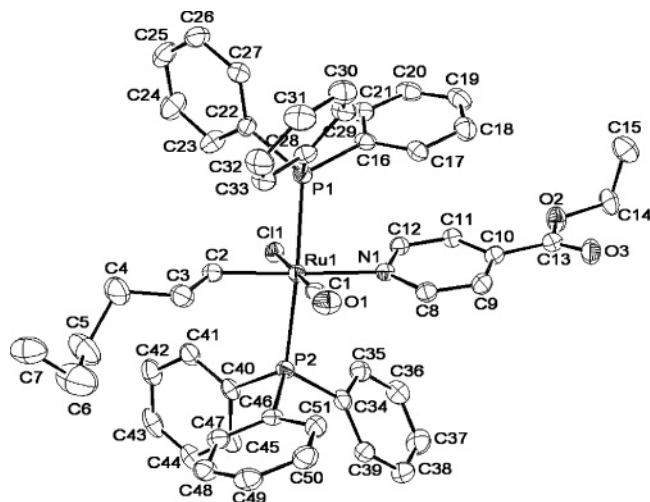


Figure 3. ORTEP representation of complex **1b** at a 50% probability level.

related osmium complex (PPh₃)₃(CO)ClOs(CH=CHPh)⁴⁷ but resembles that in the recently reported divinyl phenylene bridged diruthenium complex {(PPh₃)₃(CO)ClRu}₂(μ-CH=CH–C₆H₄–CH=CH-1,4).³² The long Ru1–P3 (Ru2–P6) bond also provides a good rationale for the facile loss and selective replacement of the phosphine ligand trans to the vinyl group.

The pyrenylvinyl ligand is nearly coplanar with the C11–Ru–C1 axis and forms a torsional angle of 178.15(4)° (C11–Ru1–C2–C3) or –165.95(5)° (C1101–Ru2–C102–C103). Again, the organic π-system and the ruthenium atoms are in direct conjugation as it is inferred from the torsional angles Ru1–C2–C3–C4 (–178.7(4)°) and Ru2–C102–C103–C104 (175.6(3)°). In contrast to **3c** the pyrenyl entities display some twist around the central C–C bonds. Thus, the vinyl-substituted and the opposing, outermost six-membered rings of the condensed π perimeter form angles of 5.9(2)° or 3.9(3)°.

Complex **1b**, the other crystallographically characterized six-coordinated derivative of this study, features a much less distorted coordination environment. A view of the molecular structure of this complex is provided as Figure 3. Angles between trans ligands span a range of 176.46(3)° to 179.46(8)° while cis angles range from 86.40(6)° to 93.45(1)°. The acute angle N1–Ru–C11 of 86.40(6)° is enforced by a hydrogen bond between an isonicotinate ortho CH proton and the adjacent chloride ligand with a C12⋯Cl1 distance of 3.153(3) Å. Such an interaction may contribute to the rotational barrier of the two isomers of HRu(CO)Cl(PPh₃)₂(isonicotinate) (vide supra). Further C⋯Cl contacts of 3.301(3) and 3.347(4) Å are observed to one ortho CH of a phenyl ring on each PPh₃ ligand. While longer, these contacts are still well below the sum of the Van der Waals radii (3.45 Å).⁵⁰ The substituted pyridine ligand is rotated by about 21° out of the plane of the equatorial donors but resides in the same plane as the vinylic double bond.

The structure of the dodecenyne complex Ru(CO)Cl(PPh₃)₂(η¹:η²-ⁿBuHC=CHC≡CⁿBu) in its crystalline state is shown as Figure S3 of the Supporting Information. σ,π-Coordination of the enynyl ligand leads to a rather long Ru–C(vinyl) σ-bond of 2.081(4) Å as compared to those in vinyl complexes like **1c**, **3c**, and **3a-P** of ca. 1.98 Å. The Ru–C bonds of 2.300(3) and 2.558(4) Å observed for the η²-coordinated alkyne part of that

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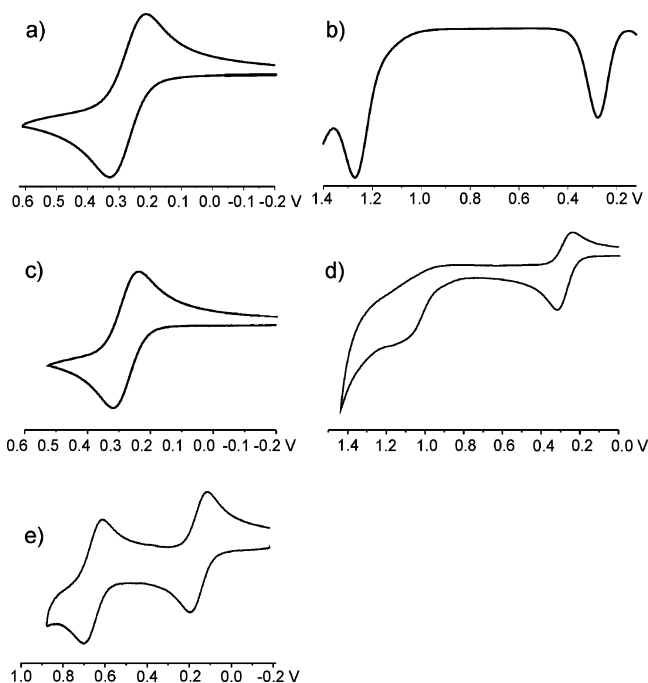


Figure 4. Cyclic voltammograms of (a) complex **1c**, (c and d) complex **2c** and (e) complex **3c** in $\text{CH}_2\text{Cl}_2/\text{NBu}_4\text{PF}_6$ (0.2 M) at $\nu = 0.1$ V/s. (b) Square wave voltammogram of complex **1c** at 10 mV/s. Potentials are given relative to the ferrocene/ferrocenium standard.

ligand are unusually long for this type of complexes,⁵¹ which points to steric crowding at the metal atom. The occupation of two cis positioned coordination sites by the enynyl ligand induces significant angular distortions in the equatorial plane. These are most evident by the acute angle C6–Ru–C8 of 66.38–(14)° spanned by the outermost carbon atoms of the enynyl ligand and the ruthenium atom and the concomitant opening of the angles C1–Ru–Cl and C1–Ru–C8 (105.68(10)° and 102.58–(15)°, respectively). These and other structural features such as the bond lengths within the enynyl part of the organic ligand (C6–C7 = 1.243(5) Å, C7–C8 = 1.418 Å, C8–C9 = 1.319–(5) Å) closely resemble those of other ruthenium enynyl complexes.⁵¹

Electrochemistry. The vinyl complexes were studied by cyclic voltammetry in $\text{CH}_2\text{Cl}_2/\text{NBu}_4\text{PF}_6$ (0.2 M) as the supporting electrolyte. Representative voltammograms of complexes **1c**, **2c**, and **3c** are shown in Figure 4; pertinent data are compiled in Table 1. A more complete set of voltammograms recorded for the studied vinyl complexes at room temperature and in a dry ice/isopropanol slush bath are provided as Figures S7–S15 of the Supporting Information. The P^iPr_3 -substituted five coordinated complexes **1c**, **2c**, and **3c** undergo a reversible one-electron oxidation at moderately positive potential. The radical cations of the five coordinated PPh_3 substituted complexes **1a**, **2a**, and **3a** are less stable as follows from $i_{p,c}/i_{p,a}$ peak current ratios of 0.37, 0.63, or 0.68 at $\nu = 0.1$ V/s at room temperature. A second one-electron oxidation occurs at substantially higher potential. This process is completely irreversible at slow sweep rates and at room temperature. For **2b,c** the chemical follow processes underlying degradation of the mono- and dications are slowed down at faster sweep rates or upon cooling the cell in a dry ice/ PrOH slush bath such that $i_{p,c}/i_{p,a}$ ratios of unity

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Table 1. Electrochemical Data of the Vinyl Complexes^a

	$E_{1/2}^{0/+}$ in V	$E_{1/2}^{+/2+}$ in V
1a	0.470 ^c	n.a.
1b	0.440 ^d	1.42 ^b
1c	0.270	1.35 ^b
2a	0.330	0.795 ^b
2b	0.360	1.05 ^b
2c	0.280	0.850 ^b
3a	0.195 ^c	0.54 ^b
	0.19 ^d	0.495 ^{c,d}
3b	0.180 ^c	0.620 ^c
	0.220 ^d	0.670 ^d
3c	0.155	0.660

^a Data in $\text{CH}_2\text{Cl}_2/0.2$ M NBu_4PF_6 at room temperature unless stated otherwise. Potentials are calibrated against the internal $\text{Cp}_2\text{Fe}^{0/+}$ couple which is set as 0.00 V. ^b Peak potential of an irreversible process at $\nu = 0.1$ V/s. ^c Half-wave potential of a partially reversible wave. ^d At 195 K.

for the first oxidation were achieved and half-wave potentials of the second oxidation process could be determined. For the 1-pyrenylvinyl complexes **3a–c** the two consecutive one-electron oxidations are more closely spaced than for the hexenyl and styryl complexes with $\Delta E_{1/2} = E_{1/2}^{+/2+} - E_{1/2}^{0/+}$ values of 350 to 500 mV. The lowering of especially the second oxidation potential renders the radical cation of **3a** and even the dication of **3c** stable, at least on the voltammetric time scale of about 10 s.

A comparison of the data in Table 1 reveals the following trends: (i) Extending the vinyl ligand's π -system from hexenyl to styryl and to 1-pyrenylvinyl shifts the oxidation potentials of both processes to more cathodic values. (ii) The effect of extending the vinyl ligand's π -system is much larger for the second oxidation. (iii) Within each series of compounds **1a–c**, **2a–c**, and **3a–c** the first oxidation potentials follow the order five-coordinate $i\text{Pr}_3$ (**1c**, **2c**, **3c**) < five-coordinate PPh_3 (**1a**, **2a**, **3a**) \approx six-coordinate isonicotinate adducts (**1b**, **2b**, **3b**). This matches the trends for the CO band shifts and the electron densities at the metal center. Similar observations have been reported for divinylphenylene-bridged diruthenium complexes $\{(\text{PR}_3)_2(\text{CO})\text{Cl}(\text{L}')\text{Ru}\}_2\{\mu\text{-C}_6\text{H}_4\text{-}(\text{CH}=\text{CH})_{2,1,3}$ or $-1,4\}$ (R = Ph, L' = 4-substituted pyridine; R = $i\text{Pr}$, L' = none).^{27,28} On the other hand, the observed cathodic shift of 40–50 mV between related complexes of series **a** and **c** is notably smaller than expected on the basis of Livers E_L parameters for PPh_3 (0.39 V) and trialkylphosphines (e.g., $\text{P}^n\text{Pr}_3 = 0.29$).⁵² In Lever's approach, the anodic oxidation potential of the metal centered $\text{Ru}^{\text{IV/III}}$ couple is expressed by eq 1, where ΣE_L denotes the sum of the E_L parameters of each ligand. Substituting two P^iPr_3 by two P^nPr_3 ligands should thus decrease the oxidation potential by about 200 mV.⁵³

$$E(\text{Ru}^{\text{IV/III}}) = 0.97 \Sigma E_L + 0.04 \text{ V} \quad (1)$$

There are two possible explanations for these trends: On one hand, the vinyl ligand becomes more electron rich as π -conjugation increases, thus decreasing its E_L parameter. This renders oxidation more facile. Support for this idea comes from a

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comparison of the calculated HOMO energies of 1-hexene (-9.97 eV),⁵⁴ propene (-9.73 eV),⁵⁵ styrene (-7.87 eV),⁵⁵ and of the 1-propylpyrenyl localized HOMO (-7.64 eV) of a pyrenyl donor/perylene-3,4,9,10-bis(dicarboximide) dyad.⁵⁶ In this view the oxidation would still be a metal-centered process. Alternatively, extending the vinyl ligand's π -system increases its contribution to the occupied FMOs that govern the oxidation processes (orbital mixing). The lower energy of the oxidized forms then mainly relies on more extensive delocalization of the positive charge(s) across the organic π -system. We will show that this is indeed the case and that replacing the hexenyl by the styryl and pyrenylvinyl ligands alters the character of the first oxidation from a metal/ligand delocalized to a ligand dominated process. The even stronger effect of vinyl ligand substitution on the second oxidation potential suggests that the alteration of relative metal/ligand contributions is even more pronounced for the second oxidation process. We also note that the first oxidation potentials of the vinyl complexes are much lower than those of comparable alkenes, styrenes, and pyrenes (cf. $E_{1/2} = 1.11$ V for α -methylstyrene and 1.20 V for styrene,⁵⁴ $E_{1/2} = \text{ca. } 0.9$ V for pyrene⁵⁷ and 2-ethynylpyrene derivatives⁴⁴), but also than those of related ruthenium complexes such as $\text{RuCl}_2(\text{CO})(\text{PPh}_3)_2(\text{py-R})$ ($\text{R} = \text{substituted pyridine}$, $E_{1/2} = 0.66\text{--}0.76$ V),⁵⁸ or $\text{HRu}(\text{CO})\text{Cl}(\text{P}^i\text{Pr}_3)_2$ ($E_p = 0.505$ V) and $\text{RuCl}_2(\text{CO})(\text{P}^i\text{Pr}_3)_2$ ($E_{1/2} = 0.625$ V) as determined by us. Cathodic shifts of ca. 400 mV for the pyrene-based oxidation have been observed in pyrenyl phenylene vinylene conjugates.⁵⁹ The low oxidation potentials observed for the vinyl complexes thus provide further evidence for an efficient conjugation within the organometallic chromophore.

IR-Spectroelectrochemistry. Owing to the synergistic nature of the metal carbonyl bond the oxidation induced shift of $\tilde{\nu}(\text{CO})$ provides an ideal tool for gauging the metal contribution to the oxidation process. A metal-centered one-electron oxidation is expected to increase the energy of the CO stretch by 100–150 cm^{-1} as less electron density is transferred from the metal atom to the π^* orbitals of the carbonyl ligand. Representative examples in ruthenium chemistry are provided by $\text{Ru}(\text{PR}_3)_2(\text{CO})_3^{0/+}$ pairs of complexes, where blue-shifts of 120–130 cm^{-1} upon oxidation have been reported.⁶⁰

With this in mind, complexes **1b–c**, **2a–c** and **3a–c** were oxidized inside an optically transparent thin layer electrolysis (OTTLE) cell⁶¹ under IR monitoring. Results of this study are displayed as Figures 5–7 and as Figures S16–S22 of the Supporting Information. Table 2 summarizes the $\tilde{\nu}(\text{CO})$, $\tilde{\nu}(\text{C}=\text{C})$, and, if applicable, the $\tilde{\nu}(\text{COOEt})$ data for the complexes in their various oxidation states.

The chemical reversibility of the second oxidations of 1-pyrenylvinyl complexes **3b,c** even allowed us to generate and

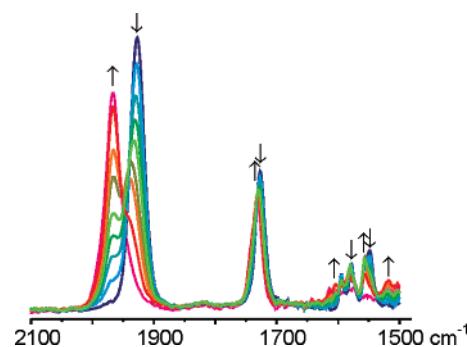


Figure 5. IR spectroelectrochemistry: spectroscopic changes during oxidation of complex **2b** ($\text{NBu}_4\text{PF}_6/\text{DCE}$).

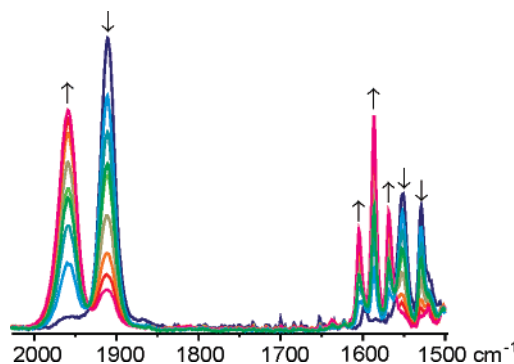


Figure 6. IR spectroelectrochemistry: spectroscopic changes during the first oxidation of complex **3c** ($\text{NBu}_4\text{PF}_6/\text{DCE}$).

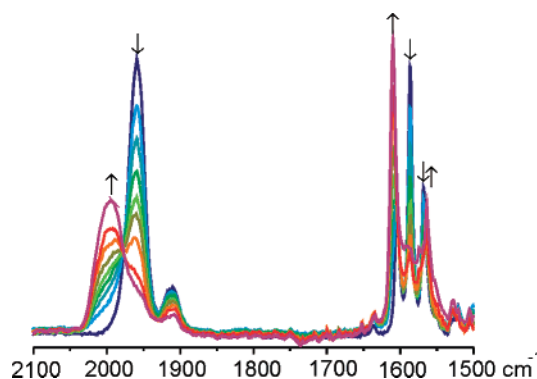


Figure 7. IR spectroelectrochemistry: spectroscopic changes during the second oxidation of complex **3c** ($\text{NBu}_4\text{PF}_6/\text{DCE}$).

to spectroscopically characterize their corresponding dications by IR spectroelectrochemistry. The second oxidation induces a further, yet even smaller blue shift of the CO band than was observed for the first oxidation. Pertinent data are provided in Table 2 (see also Figures 7 and S22 of the Supporting Information).

Comparison of the data in Table 2 reveals the following: (i) For each series **a–c** the overall CO band shift upon oxidation consistently decreases with increasing π conjugation of the substituent at the vinyl ligand. (ii) Oxidation induced CO band shifts are larger for the five-coordinate complexes and are seemingly independent of the nature of the phosphine ligands. The distinctly lower CO band shifts in the 18 VE isonicotinate adducts show that the weak π acceptor pyridine coligands serve as electron buffers to ruthenium as is indicated by the concomitant blue-shift of the isonicotinate ester band. (iii) Even for the hexenyl complexes **1b,c** the CO band shifts are considerably smaller than is expected of a metal centered

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Table 2. IR Data for Complexes **1a–c**, **2a–c**, and **3a–c** in Their Various Oxidation States

	$\tilde{\nu}$ (CO)	$\Delta\tilde{\nu}$ (CO)	$\tilde{\nu}$ (C=C)
1a	1931		1587(w)
1b	1923		1726 ^b
1b⁺	1968	45	1731 ^b
1c	1906		1586(m)
1c⁺	1983	77	1580(m)
2a	1912		1595(w), 1585(m), 1562(m), 1553(m)
2b	1927		1726(s), ^b 1594(w), 1578(m), 1548(m)
2b⁺	1967	40	1730(s), ^b 1613(w), 1580(m), 1556(m), 1519(w)
2c	1911		1595(w), 1579(m), 1554(m)
2c⁺	1976	65	1595(w), 1579(m), 1576(m), 1554(m)
3a	1935		1600(w), 1584(w), 1556(m), 1531(m)
3a⁺	1983	48	1600(sh), 1588(m)
3b	1927		1727, ^b 1599(w), 1550(m), 1527(m)
3b⁺	1962	35	1730, ^b 1601(m), 1587(m),
3b²⁺	1984	22	1741, ^b 1612(m), 1559(m), 1521(w)
3c	1911		1599(w), 1552(s), 1529(s)
3c⁺	1959	48	1605(s), 1586(s), 1569(s)
3c²⁺	1993	34	1610(s), 1574(w), 1564(m), 1529(w)

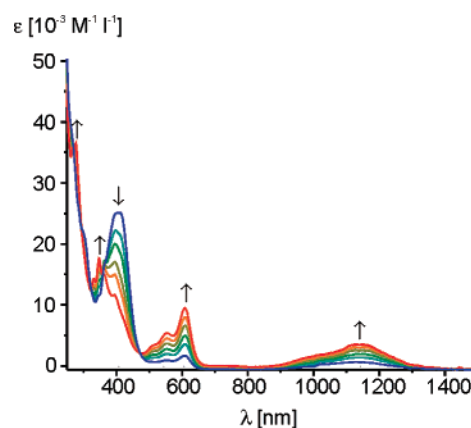
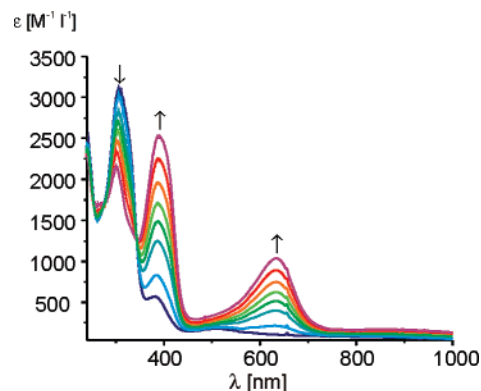
^a Energies are given for 1,2-C₂H₄Cl₂ solutions in cm⁻¹. ^b Isonicotinate ester band.

Table 3. Electronic Spectra of the Vinyl Complexes in Their Various Oxidation States

	λ_{\max} (ϵ_{\max} [M ⁻¹ cm ⁻¹])
1b	400 (2200)
1c	290(sh, 3700), 328(1120), 383(1600), 512(320)
1c⁺	296(8600), 340(sh), 508(1800), 980(200)
2a	310(14300), 376(2750), 490(820)
2b	308(23000), 406(3700), 522 (300)
2b⁺	397(3600), 686(6500), 967(595), 1020(545)
2c	272(sh, 1750), 306(3200), 381(580), 510(290)
2c⁺	272(1800), 301(2270), 388(2640), 633(1145), 875(280)
3a	296(sh, 13700), 386(12200), 486(1600), 584(1200)
3a⁺	350(sh, 8400), 363(9000), 385(9200), 465(3800), 603(2300), 1150(1700)
3b	265(sh, 39000), 285(sh, 25000), 300(23000), 414(29000), 603(420)
3b⁺	278(37000), 348(18000), 608(10000), 1139(4100)
3c	264(sh, 15700), 309(11000), 414(17000), 510(1200)
3c⁺	277(13000), 346(8800), 376 (8900), 395(8400), 415(7700), 509(3900), 551(8100), 594(11900), 704(1100), 1086(5650)

oxidation. Taking the redox pairs Ru(PR₃)₂(CO)₃^{0/+} as benchmark systems and assuming a CO band shift of 130 cm⁻¹ for a ruthenium centered process, the metal contribution to oxidation of the PⁱPr₃ complexes can be estimated as ca. 60% (**1c**), 50% (**2c**), and as ca. 35% (**3c**) of that in the benchmark systems. For complexes **3b,c**, the second oxidation is even more centered on the 1-pyrenylvinyl ligand than the first one. IR-spectroelectrochemistry thus provides clear evidence for “non-innocent” behavior of even simple alkyl substituted vinyl ligands in these ruthenium complexes.

Vis–NIR Spectroelectrochemistry. Electronic spectra of the parent hexenyl and styryl complexes in the visible region are dominated by an only moderately intense absorption band of mainly vinyl ligand to metal charge transfer (LMCT) for the five-coordinated complexes of series **a** and **c** or mixed vinyl to pyridine ligand/metal-to-pyridine ligand (LLCT/MLCT) character for complexes of series **b**. For complexes **3a–c** the pyrenyl-based absorption appears as an intense band near 410 nm without resolved vibrational structure (see Table 3 and Figure 8). We note a substantial red shift of this band with

**Figure 8.** UV–vis–NIR spectroelectrochemistry: spectroscopic changes upon the first oxidation of complex **3b** (NBu₄PF₆/DCE).**Figure 9.** UV–vis–NIR spectroelectrochemistry: spectroscopic changes upon the first oxidation of complex **2c** (NBu₄PF₆/DCE).

respect to pyrene,⁶² 1-phenylpyrene derivatives,⁵⁷ and 1-ethynylpyrene,⁶² whose optical absorptions peak at 330–350 nm. This, and the loss of the characteristic vibrational fine structure of the pyrene chromophore signal efficient electronic conjugation between the pyrenyl and the vinylruthenium entities.^{57,63} The five-coordinated derivatives **1a,c**, **2a,c** and **3a,c** also display a weak absorption ($\epsilon < 1000$ M⁻¹ cm⁻¹) in the 500–600 nm range that are typical of d–d type transitions of d⁶ ML₅ complexes of ruthenium.^{64–66}

Oxidation of the vinyl complexes in an OTTLE cell under UV–vis–NIR monitoring induces the collapse of the LMCT or LLCT/MLCT bands and the simultaneous growth of intense absorptions at lower energies in the visible and the near-infrared (NIR, see Figures 8, 9 and Figures S23–S25 of the Supporting Information). With clearly discernible vibrational splittings of 1500–1570 cm⁻¹ the overall appearance of the individual bands is highly reminiscent of that of alkene and arene radical cations, including that of styrene and various pyrene derivatives.^{67–69}

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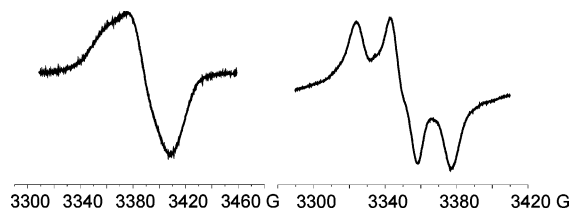


Figure 10. X-band ESR spectra of the radical cation of complex **2c** ($\text{NBu}_4\text{-PF}_6/\text{DCE}$) at 110 K (left) and at room temperature (right).

Table 4. Experimental and Calculated ESR Data of Oxidized Mononuclear Ru–Vinyl Complexes^a

	g_{iso}^b	g_{av}^c	g_1	g_2	g_3	Δg
1c ⁺		2.023	2.050	2.018	2.001	0.049
2b ⁺	2.010	2.0473	2.0653	2.0382		0.0271
2c ⁺	2.0448 ^d					
3b ⁺	2.0243	2.0242	2.039	2.0225	2.011	0.0280
3c ⁺	2.0032 ^e	2.0132	2.0263	2.0066		0.0197
1c ^{Me+}		2.0248	2.0958	2.0553	2.0452	0.0506
1c ⁺		2.0186	2.0894	2.0513	2.0460	0.0434
2b ^{Me+}		2.0116	2.0829	2.0560	2.0503	0.0326
2c ^{Me+}		2.0221	2.0606	2.0382	2.0320	0.0286
3c ^{Me+}		2.0159	2.0389	2.0271	2.0266	0.0123

^a Radical cations were electrogenerated ($\text{CH}_2\text{Cl}_2/\text{NBu}_4\text{PF}_6$) inside an ESR tube. ^b Isotropic g -value in fluid solution. ^c Average g -value at 77 K. ^d $A(^{31}\text{P}) = 21.5$ G. ^e $A(^{31}\text{P}) = 17.2$ G.

There is again a notable red-shift of the bands in the oxidized vinyl complexes when compared to the oxidized forms of their purely organic counterparts. As an example, the absorption bands of the parent styryl radical cation at 344 and 614 nm⁶⁷ are shifted to 388 and 633 nm in **2c**⁺. Photoionized trans-1-styrylpyrene absorbs at 580 nm.⁷⁰ The corresponding radical cation band of **3b,c**⁺ is observed at 608 and 594 nm, respectively. The results of UV–vis–NIR spectroelectrochemistry thus confirm the dominant organic character of the vinyl complex radical cations and extensive electron delocalization within the metalloorganic chromophore.

ESR Spectroscopy. Metal-centered paramagnetic species display fundamentally different properties in their ESR spectra than organic-based radicals. This is particularly true for Ru(III) complexes, which are mostly ESR silent at room-temperature owing to rapid relaxation processes.⁶⁰ Spectra recorded at low temperatures, usually in frozen matrices, display axial or rhombic g -tensors with large g -anisotropies, that is, large differences between the individual components of the g -tensor, and average g -values distinctly different than the free electron value g_e of 2.0023.^{71–73} Organic radicals, on the other hand, typically display richly structured isotropic signals at room temperature, g -values close to that of the free electron, and g -value anisotropies which are too small to be resolved in the X-band. ESR spectroscopy is thus ideally suited to probe the character of the vinyl complex radical cations.

Oxidized samples of most vinyl complexes were electrogenerated inside an ESR tube. With the exception of **1c**⁺ and **2b**⁺, which became ESR active only upon cooling, all other

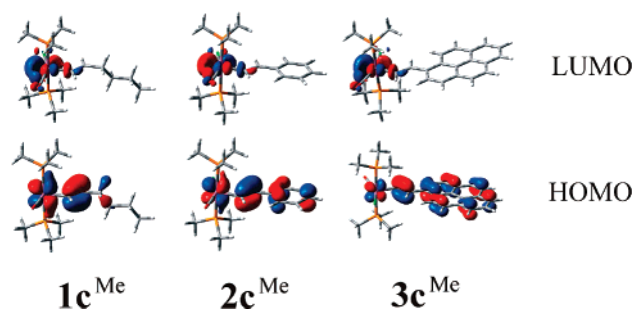
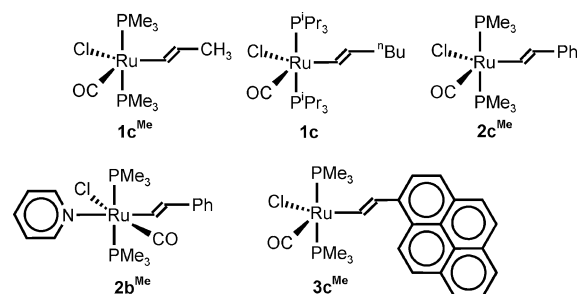


Figure 11. Contour plots of the frontier orbitals of complexes **1c**^{Me}, **2c**^{Me}, and **3c**^{Me}.

Chart 2. Vinyl Model Complexes Employed in This Study and Their Designation



investigated radical cations display intense isotropic spectra in fluid solution at room temperature (see Figure 10 and Figures S26 to S29 of the Supporting Information and Table 4). This bears clear witness to their dominant organic character. Some metal contribution to the SOMO is evidenced by the line broadening, which in most cases conceals hyperfine splittings to the vinyl protons and the P-atoms of the phosphine ligands. Exemptions are the radical cations **2c**⁺ and **3c**⁺ where resolved hyperfine couplings to the ³¹P nuclei of 21.5 and 17.2 G, respectively, were observed in fluid solution. Further manifestations of metal contribution to the SOMO orbital are the deviation of the measured g -values from g_e and some discernible axial or rhombic splitting of the g -tensor in frozen matrices (110 K). The g -value anisotropies are, however, much smaller than those usually associated with Ru(III) species.⁷⁴

Smaller g -tensor anisotropies Δg and smaller deviations of the average g -value from g_e are observed as the covalency of the Ru–ligand bond increases and more spin density is shifted to the ligand. Recent examples of Ru(III) alkynyl complexes document close correlations between (calculated) metal contributions to the SOMO and the Δg parameter.^{75,76} In none of these cases, however, g -value anisotropies as small as those in mono- and dinuclear^{27,28} vinyl complexes have been observed.

Quantum Chemistry. Quantum chemical calculations on models of the mononuclear ruthenium vinyl complexes were carried out in order to (i) assess metal versus ligand contributions to the frontier molecular orbitals (fmos) that are involved in the redox chemistry and in the low-energy optical transitions, (ii) investigate how extending π -conjugation within the vinyl ligand affects the fmo composition, (iii) rationalize the electronic spectra of the complexes, (iv) establish how oxidation changes

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Table 5. Calculated FMO Compositions of Mononuclear Ru–Vinyl Model Complexes.

model	MO	E (eV)	prevailing character	Ru	vinyl	R	CO	Cl	PR3
1c^{Me}	LUMO+1	−0.21	d _{Ru}	60	1	0	6	3	30
	LUMO	−1.62	d _{Ru}	60	13	1	11	3	11
	HOMO	−5.30	d _{Ru} + π _{vinyl}	46	43	4	0	1	2
	HOMO-1	−6.15	d _{Ru}	61	4	0	17	16	2
1c	LUMO+1	−0.16	d _{Ru}	56	0	0	7	5	32
	LUMO	−1.39	d _{Ru} + PR ₃	59	13	1	9	3	15
	HOMO	−5.38	d _{Ru} + π _{vinyl}	40	41	5	0	0	13
	HOMO-1	−6.33	d _{Ru} + Cl	63	4	0	17	14	1
2c^{Me}	LUMO+1	−0.44	π _{vinyl} + π _{Ph}	8	32	58	0	0	1
	LUMO	−1.55	d _{Ru}	60	13	1	11	3	11
	HOMO	−5.38	d _{Ru} + π _{vinyl} + π _{Ph}	28	37	30	0	0	4
	HOMO − 1	−6.55	d _{Ru}	60	5	0	17	17	1
2b^{Me}	LUMO + 3	−0.17	d _{Ru}	58	2	0	18	5	14
	LUMO + 2	−0.20	π _{vinyl} [*] + π _{Ph} [*]	8	28	61	0	0	1
	LUMO + 1	−0.40	π _{py} [*]	3	0	0	0	0	1
	LUMO	−0.98	π _{py} [*]	4	0	0	0	0	0
	HOMO	−5.15	d _{Ru} + π _{vinyl} + π _{Ph}	29	39	29	0	0	2
	HOMO − 1	−6.25	d _{Ru}	50	4	0	32	12	0
3c^{Me}	LUMO + 1	−1.54	π _{pyr} + π _{vinyl}	3	10	87	0	0	0
	LUMO	−1.59	d _{Ru}	60	12	1	12	3	12
	HOMO	−5.08	π _{pyr} + π _{vinyl} + d _{Ru}	14	21	63	0	0	2
	HOMO − 1	−6.14	π _{pyr} + π _{vinyl} + d _{Ru}	21	13	63	0	1	2

Table 6. TDDFT (G03/PBE0/CPM (CH₂Cl₂)) Singlet Excitation Energies (eV) for **1c**, **2b^{Me}**, **2c^{Me}** (G03/PBE0/CPCM (CH₂Cl₂)), and **3c^{Me}** (G03/PBE1PBE/CPCM (CH₂Cl₂)) with Oscillator Strengths Larger than 0.001

model	state	main components (%)	transition energy [eV (nm)]	osc str	expt trans [nm]	ext coeff [M ^{−1} cm ^{−1}]
1c	a ¹ A	91 (HOMO → LUMO)	2.28 (544)	0.05	512	320
	b ¹ A	88 (HOMO-2 → LUMO)	3.18 (390)	0.027	383	1600
	f ¹ A	82 (HOMO → LUMO+1)	3.60 (344)	0.024	328	1120
	h ¹ A	37 (HOMO-3 → LUMO)	4.14 (299)	0.029	290	3700
	i ¹ A	30 (HOMO-1 → LUMO+1)	4.35 (285)	0.065		
2b^{Me a}	b ¹ A	90 (HOMO → LUMO)	2.34 (529)	0.005	521	300
	c ¹ A	98 (HOMO-1 → LUMO)	3.16 (392)	0.025	394	3700
	f ¹ A	70 (HOMO-3 → LUMO)	4.05 (306)	0.025	308	23000
	g ¹ A	98 (HOMO → LUMO+1)	4.23 (293)	0.634		
2c^{Me}	b ¹ A	98 (HOMO → LUMO)	3.50 (354)	0.133	383	580
	e ¹ A	97 (HOMO → LUMO+1)	4.08 (304)	0.017	306	3200
	f ¹ A	78 (HOMO-1 → LUMO+3)	4.19 (296)	0.181		
	g ¹ A	87 (HOMO → LUMO+2)	4.26 (291)	0.464		
3c^{Me}	a ¹ A	92 (HOMO → LUMO)	2.18 (569)	0.003	510	1200
	b ¹ A	92 (HOMO → LUMO+1) (π π [*])	2.98 (416)	0.938	414	17000
	f ¹ A	80 (HOMO → LUMO+2)	3.45 (359)	0.042	n.d.	n.d.
	h ¹ A	37 (HOMO-1 → LUMO+1)	3.92 (316)	0.077	309	11000
		40 (HOMO → LUMO+4)				
	i ¹ A	30 (HOMO-1 → LUMO+1)	4.01 (309)	0.124		

^a Contributions missing to 100% correspond to py ligand-based orbitals.

the bonding within the organometallic Ru-vinyl chromophore, and (v) compare the experimentally observed CO and C=C frequencies for the accessible oxidation states and the *g*-parameters of the radical cations to those predicted by theory. The phosphine ligands were modeled as PMe₃ for most calculations. To probe for the effect of such simplification, the model complex Cl(CO)(PMe₃)₂Ru(CH=CHMe) (**1c^{Me}**) was compared to fully optimized **1c**. The effect of six versus five coordination was finally addressed by comparing Cl(CO)-(PMe₃)₂Ru(CH=CHPh) (**2c^{Me}**) and (py)Cl(CO)(PMe₃)₂Ru(CH=CHPh) (**2b^{Me}**) as models of complexes **2a,c** and **2b**. The model complexes employed in the computational study and their designation are summarized in Chart 2.

The optimized structures of the model complexes agree well with the experimental ones (see Tables S11 to S14 of the Supporting Information). Deviations between experimental and calculated structures are restricted to the immediate vicinity of the vinyl bond.

Contour plots of crucial fmos of the complexes **1c^{Me}**, **2c^{Me}**, and **3c^{Me}** are provided as Figure 11, while Table 5 lists their compositions as obtained by Mulliken analysis. These data fully conform to our experimental observation that metal contribution to the occupied frontier levels decreases as the π-conjugation of the vinyl ligand increases. For the propenyl and hexenyl systems **1c^{Me}** and **1c** the HOMO is fully delocalized over the Ru-vinyl entity. Vinyl ligand contributions to the HOMO become more and more important as the organic π-system extends. It increases from 47% in **1c^{Me}** or 46% in fully optimized **1c** to 67% in **2c^{Me}** and 84% in **3c^{Me}**. This clearly justifies denoting them as “non-innocent”. A comparison of model complex **1c^{Me}** and the accurate model of **1c** shows, that approximating PⁱPr₃ by PMe₃ coligands underestimates the phosphine contribution to the HOMO and LUMO orbitals. The LUMO of all five coordinated complexes is essentially a ruthenium d-orbital with contributions from the phosphine and carbonyl ligands such that the electronic transition at the lowest

Table 7. Calculated IR Parameters of Vinyl Complexes in Their Various Oxidation States.

		$n=0$	$n=1$	$n=2$
1c ^{Me n+}	$\tilde{\nu}$ (CO)	1926.9	1988.5	2045.0
1c ⁿ⁺	$\tilde{\nu}$ (CO)	1924.5	1986.0	2037.5
2b ^{Me n+}	$\tilde{\nu}$ (CO)	1916.0	1961.0	2001.5
	$\tilde{\nu}$ (CC)	1554.2	1597.0	1607.4
2c ^{Me n+}	$\tilde{\nu}$ (CO)	1927.6	1978.9	2024.9
	$\tilde{\nu}$ (CC)	1549.9	1595.7	1607.9
3c ^{Me n+}	$\tilde{\nu}$ (CO)	1926.9	1967.2	2002.5
	$\tilde{\nu}$ 1(CC)	1550.2	1606.7	1608.4
	$\tilde{\nu}$ 2(CC)	1526.6	1584.0	1576.8

energy is assigned as a mixed ligand-to-metal/ligand-to-ligand charge-transfer band (LMCT/LLCT). This is also borne out by TDDFT calculations which account for CH₂Cl₂ solvation. As is seen from the comparison in Table 6, the calculated spectra reproduce the experimental ones well. On the basis of these calculations, the lowest energy band at ca. 510 nm is assigned as the HOMO–LUMO transition. Higher energy transitions within the styryl and pyrenyl complexes have dominant $\pi \rightarrow \pi^*$ character owing to sizable contributions of the vinyl ligand to the LUMO + n and the HOMO – n orbitals, particularly for the pyrenyl complex. For six-coordinated **2b**^{Me}, however, two π^* orbitals of the pyridine coligands are interspersed between the HOMO and the appropriate unoccupied metal d orbital such that the two lowest electron transitions assume mixed arylvinyl-to-pyridine (LLCT) and metal-to-pyridine charge transfer (MLCT) character.

To trace the effect of successive charge removal from the occupied frontier levels diagnostic IR parameters such as $\tilde{\nu}$ (CO) and $\tilde{\nu}$ (C=C) of the model complexes were calculated for the neutral compounds and the oxidized mono- and dications. The compilation in Table 7 shows the expected decrease of oxidation-induced CO band shifts with decreasing HOMO metal character, paralleling our experimental results. Thus, for the first oxidation, $\Delta\tilde{\nu}$ (CO) diminishes from 61 cm⁻¹ in the propenyl complex **1c**^{Me} to 51 in **2c**^{Me} and 40 cm⁻¹ for **3c**^{Me}. The calculations slightly underestimate $\Delta\tilde{\nu}$ (CO), which is mostly due to somewhat high $\tilde{\nu}$ (CO) values for the neutral complexes. For the styryl and pyrenylvinyl complexes the energies of the vinyl and substituent C=C bands are predicted to increase with the overall oxidation state. We note again good agreement with the experimental data, particularly for the pyrenylvinyl complexes where these vibrations are rather intense. Computed ESR parameters as they are given in Table 4 fully reproduce the experimentally observed small g -anisotropies Δg , the systematic decrease of Δg with increasing π -conjugation and lower metal character of the SOMO as well as the small deviations of the average g -values from g_e .

Conclusions and Perspectives. Taken together, our experimental and computational results provide compelling evidence that even simple alkenyls behave as “non-innocent” ligands when coordinated to Ru(CO)Cl(PR₃)₂(L) (L = electron donating or weakly π -accepting coligand) entities. Extending the conjugation within the vinyl ligand gradually alters the nature of the HOMO and, accordingly, of the anodic redox processes from fully delocalized across the metal–vinyl entity to ligand dominated with smaller, yet still discernible metal contributions. Such a situation is opposite to that encountered in most π -bridged dimetal complexes that are often invoked as models

and principal constituents of organometallic “molecular wires”.^{77–81} There, the occupied frontier levels are metal centered and the metal atoms interact through direct orbital overlap with the π -conjugated bridge. In our case, delocalization of the ligand centered HOMO onto the metal atoms suggests another possible way to construct extended electron conducting systems from vinyl ruthenium building blocks by employing the metal atoms as the conduit between two redox active ligands. Such a possibility has already been realized in a mononuclear bis-(alkynyl) platinum complex.⁸² Our results add to a growing body of redox-active organometallic compounds where the organic π -ligand heavily participates in electron-transfer processes.^{79,80,83–85} Five-coordinated mononuclear ruthenium complexes as they were investigated herein offer an additional vacant coordination site trans to the vinyl ligand. This allows for the easy assembly of di- and oligonuclear systems with possibly high conductivity along the vinyl–metal-bridge chain by interconnecting them through appropriate bridging ligands or by employing alkynes with an additional coordinating functionality in the hydorruthenation reaction. This is the subject of ongoing research in our laboratories.

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Supporting Information Available: Synthetic procedures and characterization of compounds, Figures displaying the crystallographically determined structures of the complexes **1b,c**, Ru(CO)Cl(PPh₃)₂(η^1 : η^2 -ⁿBuHC=CHC≡CⁿBu), **3a-P**·**2CH₂Cl₂** and **3c**·**CHCl₃** (Figures S1–S5); tables detailing the crystal data and structure refinement (Tables S1–S8); voltammograms of complexes **1a–c**, **2a–c** and **3a–c** at room temperature and in an isopropanol/dry ice slush bath (Figures S6–S14); IR-spectroscopic changes upon the first oxidation of complexes **1b**, **1c**, **2b**, **3a**, and **3b** (S15–S20), and upon the second oxidation of complex **3b** (S21); Changes in the UV–vis–NIR spectrum of **1b** (S22), **2b** (S23), and **3b** (S24); ESR spectra of the radical cations of complexes **1c**, **3a,b**, and **3b,c** (S25–S28); tables of the G03/B3LYP and ADF/BP calculated bond parameters for the fully optimized complexes **1c** and **3c**^{Me} (Tables S9, S10) and contour plots of all orbitals involved in the electronic transitions of Table 6. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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